

The Synthesis and Characterisation, including X-ray Diffraction Study, of $[\text{Th}\{\eta\text{-C}_5\text{H}_3(\text{SiMe}_3)_2\}_3]$; the First Thorium(III) Crystal Structure†

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Reduction of $[\text{ThCp}''_2\text{Cl}_2]$ [$\text{Cp}'' = \eta\text{-C}_5\text{H}_3(\text{SiMe}_3)_2$] with Na–K in PhMe yields $[\text{ThCp}''_3]$, which is monomeric in the solid state, as determined by X-ray crystallography, and is the first crystalline tri- η -cyclopentadienylmetal complex, $\langle \text{Th-C}(\eta\text{-}) \rangle 2.80(2)$ Å.

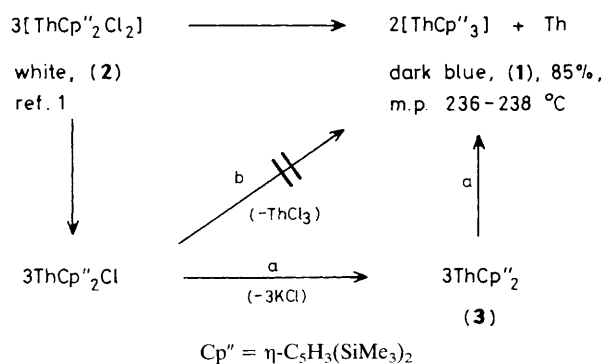
We report the synthesis and X-ray characterisation of the exceedingly air- and moisture-sensitive but thermally stable, hydrocarbon-soluble, blue, paramagnetic, crystalline f¹ complex $[\text{ThCp}''_3]$ [$\text{Cp}'' = \eta\text{-C}_5\text{H}_3(\text{SiMe}_3)_2$] (**1**). As far as we are aware, there are no previous X-ray single-crystal structural data either for a Th^{III} complex or a homoleptic tris(η -cyclopentadienyl)metal complex.

Complex (**1**) was obtained by an unexpected reductive disproportionation from the colourless f⁰ complex $[\text{ThCp}''_2\text{Cl}_2]$ (**2**)¹ and a five-fold excess of Na–K alloy in toluene. We believe that (**1**) is formed (pathway a in Scheme 1) via the transient thorocene (**3**), because (i) the yield of (**1**) was virtually quantitative in Cp'' and Th metal was a coproduct; (ii) metathesis (pathway b in Scheme 1) via $\text{ThCp}''_2\text{Cl}$ is unlikely, *cf.*,² the stable $[\{\text{UCp}''_2(\mu\text{-Cl})\}_2]$ (**4**); (iii) we find that (**4**) is reduced by Na–K alloy in PhMe to U; and (iv) there is trapping evidence both for UCp''_2 ($\text{Cp}'' = \eta\text{-C}_5\text{Me}_5$)³ and ThCp''_2 ⁴ {*e.g.*, $[\text{ThCp}''_2\text{Cl}_2] + \text{Mg}(\text{buta-1,3-diene}) \rightarrow [\text{ThCp}''_2(\eta\text{-C}_4\text{H}_6)]$ }.

It is evident that the thorocene(IV) chloride (**2**) is difficult to reduce. Thus, it did not react with Na–Hg in toluene upon prolonged reflux, although³ $[\text{UCp}''_2\text{Cl}_2]$ and Na–Hg in tetra-

hydrofuran (thf) gave $[\text{UCp}''_2(\mu\text{-Cl})_2\text{Na}(\text{thf})_2]$. Furthermore,⁵ $[\text{ThCp}''_2\text{Cl}_2]$ did not undergo electrochemical reduction in tetrahydrofuran (thf) or MeCN below -2.7 V [*cf.*, for the U analogue, $-E_3 = 1.30$ (MeCN) or 1.22 (thf)].

Complex (**1**) is monomeric in the crystal, with the three equivalent $\eta\text{-C}_5\text{H}_3(\text{SiMe}_3)_2$ groups distributed symmetrically



Scheme 1. Reagents and conditions: Na–K, PhMe, 20 °C, 24 h. Compound (**1**) was characterised by microanalysis, ¹H n.m.r. spectroscopy [a broad single peak (ω_1 900 Hz) centred at 5.3 p.p.m. attributed to SiMe₃ protons], and X-ray diffraction (Figure 1).

† No reprints available.

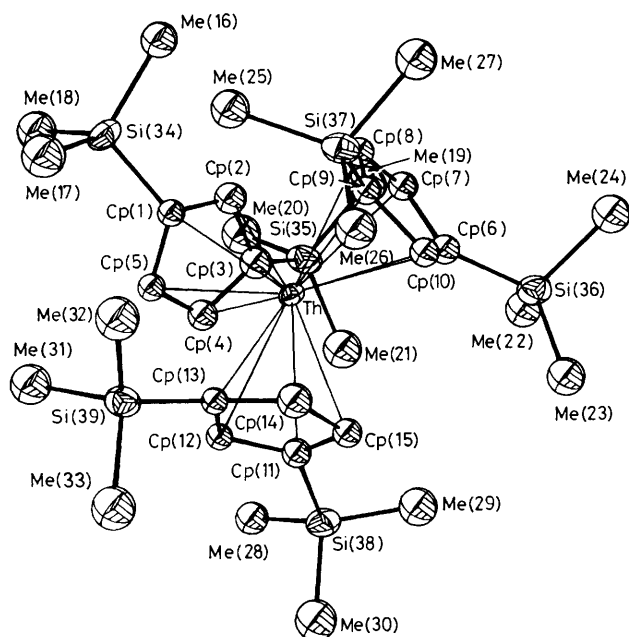


Figure 1. The molecular structure, and atom labelling scheme, for $[\text{ThCp}^*_3]$ (**1**). Relevant dimensions are: Th–Cp(1) aver. 2.80(2), Th–Cp(2) aver. 2.81(2), Th–Cp(3) aver. 2.79(2), Th–cent.(1) 2.519, Th–cent.(2) 2.526, Th–cent.(3) 2.509 Å; cent.(1)–Th–cent.(2) 121.11(4), cent.(1)–Th–cent.(3) 118.86(3), cent.(2)–Th–cent.(3) 120.03(4)°. [Cent.(1), cent.(2), and cent.(3) are the centroids of the cyclopentadienyl rings.]

about the quasi-nine-co-ordinate thorium atom, Figure 1.† The average Th–C(η -) distance of 2.80(2) Å in $[\text{ThCp}^*_3]$ (**1**) is close to that in $[\text{ThCp}^*_2\text{Cl}_2]$ (**2**), 2.78(1) Å.¹ The co-ordination environment around Th can alternatively be described as trigonal planar, the average of the three cent.–Th–cent. angles

† Crystal data for $[\text{ThCp}^*_3]$ (**1**): $M = 860.42$, space group $P2_1/c$, $a = 18.006(5)$, $b = 13.783(5)$, $c = 19.359(8)$ Å, $\beta = 112.83(5)^\circ$, $U = 4427.95$ Å³, $Z = 4$, $D_c = 1.29$ g cm⁻³, $\mu = 45.52$ cm⁻¹ (Mo- K_α).

The structure of complex (**1**) was solved by routine heavy atom methods and refined to $R = 0.049$, $R' = 0.067$ for 2224 observed reflections, measured on a CAD-4 diffractometer. Details of the data collection process have previously been given.⁶

Atomic co-ordinates, bond lengths and angles, and thermal parameters have been deposited at the Cambridge Crystallographic Data Centre. See Notice to Authors, Issue No. 1, 1986.

$= 120^\circ$ (cent. denotes the centroid of each $\bar{\text{Cp}}^*$ ligand); $\langle \text{Th-cent.} \rangle 2.52$ Å [*cf.*, 2.51 Å in (**2**)].

The organometallic chemistry of the actinides, and especially of U and Th, is developing rapidly, mainly for complexes in oxidation state +4.⁷ The +3 state is also readily accessible for U, and X-ray data for several compounds are available, including $[\{\text{UCp}^*_2(\mu\text{-Cl})\}_3]$,³ $[\{\text{UCp}^*_2(\mu\text{-X})\}_2]$ ($X = \text{Cl}$ or Br),² $[\text{UCp}^*_2\text{H}(\text{dmpe})]$ ($\text{dmpe} = \text{Me}_2\text{P}[\text{CH}_2]_2\text{PMe}_2$),⁸ $[\text{UCp}_3(\text{thf})]$,⁹ $[\text{U}(\eta\text{-indenyl})_3]$,¹⁰ $[\text{U}(\eta\text{-C}_5\text{H}_4\text{SiMe}_3)_3(\text{CNEt})]$,¹¹ and $[\text{UCp}_3(\text{Bu}^n)]^-$.¹² Thorium(III) complexes are very rare, but the hydrocarbon-insoluble compounds ThCp_3 , $\text{Th}(\text{C}_5\text{H}_4\text{Me})_3$, and $\text{Th}(\text{indenyl})_3$ have been reported.^{7,13}

We thank S.E.R.C. and A.E.R.E. Harwell for the award of a C.A.S.E. studentship to P. C. B., N.S.F. for partial support for J. L. A., and Dr. D. Brown for his interest.

Received, 7th April 1986; Com. 453

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